



**NATIONAL INSTITUTE OF TECHNOLOGY,
TIRUCHIRAPPALLI**

DEPARTMENT OF CHEMISTRY

COURSE PLAN - PART I			
Name of the programme and specialization	B.Tech. I (Civil Engineering)		
Course Title	Chemistry		
Course Code	CHIR11	No. of Credits	3
Course Code of Pre-requisite subject(s)	Nil		
Session	July 2021	Section (if, applicable)	B
Name of Faculty	Dr. Sunandan Sarkar	Department	Chemistry
Official Email	ssarkar@nitt.edu	Telephone No.	9153484492
Name of Course Coordinator(s) (if, applicable)			
Official E-mail		Telephone No.	
Course Type (please tick appropriately)	<input checked="" type="checkbox"/> Core course	<input type="checkbox"/> Elective course	
Syllabus (approved in BoS)			
<p>Electrochemistry and Corrosion Cell EMF- its measurement and applications - concentration cell - electrode electrolyte concentration cell - concentration cell with and without transference - Dry corrosion and wet corrosion, mechanisms, types of corrosion, Differential metal corrosion, differential aeration corrosion, intergranular, Passivity, Pitting, Polarization - Chemical conversion coatings and organic coatings- Paints, enamels.</p> <p>Phase rule Definition of terms – phase- components- degree of freedom- derivation of Gibbs phase rule – one component system – H₂O, CO₂, Sulfur – Two-component system – Eutectic systems – reduced phase rule - Pb-Ag system – Compound Formation with congruent melting – Zn- Mg Alloy system- Copper-nickel alloy system - systems with incongruent melting – Na₂SO₄- H₂O system and simple three-component systems.</p> <p>Water Sources, Hard & soft water, Estimation of hardness by EDTA method, Scale & Sludge- Caustic embrittlement - softening of water, zeolite process & demineralization by ion exchangers, boiler feed water, internal treatment methods-specifications for drinking water, BIS & WHO standards, treatment of water for domestic use, desalination - Reverse osmosis & Electrodialysis.</p> <p>Spectroscopy Interaction of electromagnetic radiation with matter, Electronic spectroscopy - Theory of electronic transitions, instrumentation, Beers Lambert law, Woodward FIESER rule, applications. IR spectroscopy - Fundamentals, Instrumentation, and applications, Raman</p>			



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spectroscopy – Fundamentals and applications.

Polymers and Composites

Concept of macromolecules- Tacticity- Classification of Polymers- Types of Polymerization Mechanism- - Ziegler Natta Polymerization - Effect of Polymer structure on properties - Important addition and condensation polymers –synthesis and properties – Molecular mass determination of polymers- Static and dynamic methods, Light scattering- Rubbers –Vulcanization – Synthetic rubbers – Conducting polymers- Composite materials.

References & Text Books

1. P.C. Jain, M. Jain, Engineering Chemistry, Dhanpat Rai Publishing Company, New Delhi, 2005.
2. P. Atkins, J.D. Paula, Physical Chemistry, Oxford University Press, 2002.
3. B.R. Puri, L.R. Sharma, M.S. Pathania, Principles of Physical Chemistry, Vishal Publishing Company, 2008
4. F.W. Billmeyer, Textbook of Polymer Science, 3rd Edition, Wiley. N.Y. 1991.
5. S.S. Darer, S.S. Umare, A Text Book of Engineering Chemistry, S. Chand Publishing, 2011.

COURSE OBJECTIVES

To introduce the student's basic principles of Electrochemistry and Corrosion. They will be familiar with phase rule & its applications. Students will know about the essential requirements of water and its importance in day-to-day life. To provide students with a brief outline of the types and applications of polymers. Finally, students will be equipped with the usage of spectroscopy in potential applications

MAPPING OF COs with POs

Course Outcomes	Programme Outcomes (PO) (Enter Numbers only)
At the end, the students should be able to:	
1. Students will learn about the Electrochemistry and phase rule.	
2. They will be familiarized with the importance of polymer and its application in industries.	
3. Additionally, a brief introduction in the area of water, spectroscopy will be very useful for the students in future endeavour.	

COURSE PLAN – PART II

COURSE OVERVIEW

This course is offered to first year B.Tech. (Civil Engineering) students. This is a 3 credit theory course and three theory classes will be conducted per week. If required, extra classes will be arranged.

COURSE TEACHING AND LEARNING ACTIVITIES

(Add more

rows)

S.No.	Week/Contact Hours	Topic	Mode of Delivery
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1	IV week of December	Electrochemistry and Corrosion Basic of Electrochemistry.	Online
2	V week of December	Cell EMF- its measurement and applications	Online
3	I week of January	Concentration cell - electrode electrolyte concentration cell - concentration cell with and without transference	
4	II week of January	Dry corrosion and wet corrosion, mechanisms, types of corrosion, Differential metal corrosion, differential aeration corrosion intergranular, Passivity, Pitting, Polarization - Chemical conversion coatings and organic coatings- Paints, enamels.	Online
5	III week of January	Phase rule Definition of terms – phase- components- degree of freedom- derivation of Gibbs phase rule – one component system – H ₂ O, CO ₂ , Sulfur, Two-component system, Eutectic systems – reduced phase rule - Pb-Ag system	Online
6	IV week of January	Compound Formation with congruent melting – Zn-Mg, Alloy system - Copper-nickel alloy system - systems with incongruent melting – Na ₂ SO ₄ - H ₂ O system and simple three-component systems.	Online
7	I week of February	Polymers and Composites Concept of macromolecules- Tacticity- Classification of Polymers Types of Polymerization Mechanism- - Ziegler Natta Polymerization - Effect of Polymer structure on properties	Online
8	II week of February	Important addition and condensation polymers – synthesis and properties. Molecular mass determination of polymers- Static and dynamic methods, Light scattering- Rubbers – Vulcanization – Synthetic rubbers – Conducting polymers- Composite materials.	Online
9	III week of February	Spectroscopy Interaction of electromagnetic radiation with matter, Electronic spectroscopy - Theory of electronic transitions, instrumentation, Beers Lambert law,	Online
10	IV week of February	Woodward FIESER rule, applications. IR spectroscopy - Fundamentals, Instrumentation, and applications, Raman spectroscopy – Fundamentals and applications.	Online
11	I week of March	Water Sources, Hard & soft water, Estimation of hardness by EDTA method, Scale & Sludge- Caustic embrittlement - softening of water,	Online
12	II week of March	zeolite process & demineralization by ion exchangers, boiler feed water, internal treatment	Online



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		methods-specifications for drinking water, BIS & WHO standards, treatment of water for domestic use, desalination - Reverse osmosis & Electrodialysis.	
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COURSE ASSESSMENT METHODS (shall range from 4 to 6)

S.No.	Mode of Assessment	Week/Date	Duration	% Weightage
1	Test I	III week of January	1hr. 30min.	30
2	Test II	III week of February	1hr. 30min.	30
3	Assignment/Viva	I week of March	4 days	10
CPA	Compensation Assessment*	II week of March	1hr. 30min.	30
6	Final Assessment *	III/IV week of March	2 hrs.	30

***mandatory; refer to guidelines on page 4**

COURSE EXIT SURVEY (mention the ways in which the feedback about the course shall be assessed)

1. Feedback from students during class committee meetings.
2. Anonymous feedback through questionnaire at the end of the semester.

COURSE POLICY (including compensation assessment to be specified)

MODE OF CORRESPONDENCE (email/ phone etc)

ssarkar@nitt.edu; Tel. No.:9153484492

COMPENSATION ASSESSMENT POLICY

- 1. This assessment is for those students who missed Test I/II due to genuine reasons.**
- 2. Compensation assessment will be conducted during the I week of March.**

ATTENDANCE POLICY (A uniform attendance policy as specified below shall be followed)

- At least 75% attendance in each course is mandatory.
- A maximum of 10% shall be allowed under On Duty (OD) category.
- Students with less than 65% of attendance shall be prevented from writing the final assessment and shall be awarded 'V' grade.

ACADEMIC DISHONESTY & PLAGIARISM



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- Possessing a mobile phone, carrying bits of paper, talking to other students, copying from others during an assessment will be treated as punishable dishonesty.
- Zero mark to be awarded for the offenders. For copying from another student, both students get the same penalty of zero mark.
- The departmental disciplinary committee including the course faculty member, PAC chairperson and the HoD, as members shall verify the facts of the malpractice and award the punishment if the student is found guilty. The report shall be submitted to the Academic office.
- The above policy against academic dishonesty shall be applicable for all the programmes.

ADDITIONAL INFORMATION, IF ANY

FOR APPROVAL

S. Sarkar 23/12/2021
Dr. Sunandan Sarkar
Course Faculty _____ CC- Chairperson _____ HOD _____

R. Sankar
Head
Department of Civil Engineering
National Institute of Technology
Tiruchirappalli - 620 015

G. S. S.



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Guidelines

- a) The number of assessments for any theory course shall range from 4 to 6.
- b) Every theory course shall have a final assessment on the entire syllabus with at least 30% weightage.
- c) One compensation assessment for absentees in assessments (other than final assessment) is mandatory. Only genuine cases of absence shall be considered.
- d) The passing minimum shall be as per the regulations.

B.Tech. Admitted in				P.G.
2018	2017	2016	2015	
35% or (Class average/2) whichever is greater.		(Peak/3) or (Class Average/2) whichever is lower		40%

- e) Attendance policy and the policy on academic dishonesty & plagiarism by students are uniform for all the courses.
- f) Absolute grading policy shall be incorporated if the number of students per course is less than 10.
- g) Necessary care shall be taken to ensure that the course plan is reasonable and is objective.